

Hydrated Crystal Lab Answers

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The crystals change form, and sometimes color, as the water is driven off. This suggests that water was present as part of the crystal structure. Such compounds are called hydrates. A hydrate that has lost its water is called an anhydrous salt. For a hydrate, the number of moles of water present per mole of salt is usually some simple, whole number.

Hydrate lab calculations - YouTube

Introduction: A hydrated crystal is a typically an ionic compound that has water molecules trapped within the crystal. This usually occurs because the compounds are hygroscopic; which means they...

Hydrated Crystal Lab Answers

5. Using your answers to numbers 2 and 4, determine the formula of the hydrate. Conclusions: 1. A hydrate is a crystal with ____ molecules adhering to the ions or molecules. 2. Define "anhydrous." 3. Write the formula for hydrated zinc sulfate with 7 molecules of water adhering to each formula unit of zinc sulfate.

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Answer to: What is a hydrated crystal? By signing up, you'll get thousands of step-by-step solutions to your homework questions. You can also ask...

CHEMISTRY LAB: HYDRATED CRYSTALS - Kwanga.net

Digication ePortfolio :: General Chemistry (Alexander Antonopoulos) by Alexander P. Antonopoulos at Salve Regina University. Lab 2: Determine the Percentage of Water in a Hydrate: The goal of this experiment is to learn how to properly calculate the ratio of salt to water, in a hydrated salt, and to calculate the percentage of water (by mass) within a hydrated salt.

Lab: Episode 703—Datasheet Name

4 is the right substance for the moisture absorber and indicator, you will need to examine the hydrated and anhydrous forms of the compound and determine the following:

- the empirical formula of the hydrate, including its water of crystallization,
- if the compound is useful as an indicator when it changes from the hydrated to the anhydrous form, and
- the mass of water absorbed by the 25 g of anhydrous compound, which the company proposes to use.

LAB: Percent Composition of Hydrated Crystals

The formula for the hydrated crystal tells how many water molecules are attached to each formula unit of crystal. For example, hydrated calcium nitrate is written: $\text{Ca}(\text{NO}_3)_2 \cdot 4 \text{H}_2\text{O}$ and is called calcium nitrate tetrahydrate. The dot in the formula shows that for every mole of calcium nitrate in the hydrated crystal, there are 4 moles of water.

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hydrate lab

Percent Composition of Hydrates

These compounds are referred to as being “hydrated” (meaning, to contain water) and are said to have a “water of hydration”. These water molecules are

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scattered within the crystal structure rather loosely; therefore heating over a low heat will drive off the water molecules.

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Hydrate Lab - Google Docs

Introduction:A hydrated crystal is a typically an ionic compound that has water molecules trapped within the crystal.This usually occurs because the compounds are hygroscopic;which means they...

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To determine the percentage of water in a hydrate. When copper (II) sulfate hydrate, a blue crystalline solid containing embedded water molecules (called a hydrate), is heated in air, it loses the water molecules and the blue solid is transformed to a white anhydrous (no water) crystal known as copper (II) sulfate.

Formula of a Hydrate Lab

$4 \cdot 2H_2O \cdot CaSO_4(s) + 2 H_2O(g)$ Therefore, you can determine the percentage of water in a hydrate by determining the mass lost. (amount of water driven off) when a known mass of hydrate is heated.

Percentage of water = $\frac{\text{Mass of water lost}}{\text{Mass of hydrate}} \times 100$.

Virtual Lab: Hydrates

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EXPERIMENT 7: HYDRATES Introduction: Background

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CRYSTALS CHEMISTRY LAB HELP - Yahoo Answers 5. Using your answers to numbers 2 and 4, determine the formula of the hydrate. Conclusions: 1. A hydrate is a crystal with ____ molecules adhering to the ions or molecules. 2. Define "anhydrous." 3. Write the formula for hydrated zinc sulfate

Bing: Hydrated Crystal Lab Answers

the crystal structure. Crystalline compounds that retain water during evaporation are referred to as being hydrated or are said to contain water of hydration. The ratio of moles of water to moles of compound is a small whole number. Example: $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ Mole ratio: 1:2 Name: barium chloride dihydrate THIS LAB: $\text{CuSO}_4 \cdot __ \text{H}_2\text{O}$

HYDRATED COMPOUNDS LAB

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HYDRATED CRYSTALS CHEMISTRY LAB HELP - Yahoo Answers

Your Hydrated Crystal formula: $\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$. Complete the following calculations. 1) Calculate the total molar mass of the formula (remember to use the number in front of the H_2O). 2) Calculate the molar mass of just the water in the formula (Use the multiple in front of the H_2O). 3) Calculate the % composition of water in the

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Hydrate Lab. Kimberly Graziano & Hyunjae Kim.
Purpose. Experimental Question: How can we experimentally determine the formula of an unknown hydrate, A? Testable Prediction: Our unknown hydrate may be a hydrate of copper(II) sulfate, magnesium sulfate, iron(III) chloride, or iron(III) nitrate. Observing our nitrate, it has a white crystalline structure, representing that similar to table salt.

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